Laboratory No. 7845 Nichols FLA. Nov 1917

CERTIFICATE OF ANALYSIS

Of Sample of ~~Phosphate~~ Rock Received from Mr Paul C Taylor

Marked #1

Sample by Date Received Nov 1-17

ANALYSIS OF DRY SAMPLE

(Dryed at 100 Deg. C.)

Phosphoric Acid	2.94	per cent.
Equi. Bone Phosphate of Lime	6.42	" "
Oxide of Iron and Alumina	1.10	" "
Insoluble Siliceous Matter	25.31	" "
Moisture, determined on ~~dry~~ sample (at 100 Deg. C.)	0.44	" "
Bone Phosphate of Lime, calculated on moist sample	6.40	" "

E. T. Carter Chemist

Laboratory No. 7854 Nichols FLA., Nov 191 2

CERTIFICATE OF ANALYSIS

Of Sample of [redacted] Rock Received from Mr. Paul C. Taylor

Marked 4 2

Sample by _____ Date Received Nov 1-12

ANALYSIS OF DRY SAMPLE

(Dryed at 100 Deg. C.)

Phosphoric Acid	1.01	per cent.
Equi. Bone Phosphate of Lime	2.22	" "
Oxide of Iron and Alumina	0.50	" "
Insoluble Siliceous Matter	9.27	" "
Moisture, determined on ~~coarse~~ sample (at 100 Deg. C.)	0.38	" "
Bone Phosphate of Lime, calculated on moist sample	2.21	" "

E. T. Cooke Chemist

Laboratory No. 7355 Nichols FLA. Nov 2 191 7

CERTIFICATE OF ANALYSIS

Of Sample of ~~Phosphate~~ Rock Received from W. Paul C. Taylor

Marked # 3

Sample by _____ Date Received Nov 1-17

ANALYSIS OF DRY SAMPLE

(Dryed at 100 Deg. C.)

Phosphoric Acid	0.66	per cent.
Equi. Bone Phosphate of Lime	1.44	" "
Oxide of Iron and Alumina	0.72	" "
Insoluble Siliceous Matter	7.88	" "
Moisture, determined on ~~coarse~~ sample (at 100 Deg. C.)	0.54	" "
Bone Phosphate of Lime, calculated on moist sample	1.43	" "

E. T. Cooke Chemist

Laboratory No. 7856 Nichols FLA., Nov 191

CERTIFICATE OF ANALYSIS

Of Sample of ~~Phosphate~~ Rock Received from Mr Burl C. Taylor

Marked #24

Sample by Date Received Nov 1 13

ANALYSIS OF DRY SAMPLE

(Dryed at 100 Deg. C.)

Phosphoric Acid	1.22	per cent.
Equi. Bone Phosphate of Lime	2.67	" "
Oxide of Iron and Alumina	0.82	" "
Insoluble Siliceous Matter	16.06	" "
Moisture, determined on ~~coarse~~ sample (at 100 Deg. C.)	0.42	" "
Bone Phosphate of Lime, calculated on moist sample	2.66	" "

E. T. Barbee Chemist

Laboratory No. 7557 Nichols FLA. Nov 7 191 7

CERTIFICATE OF ANALYSIS

Of Sample of ~~Phosphate~~ Rock Received from Nichols

Marked #5

Sample by Date Received Nov 14-17

ANALYSIS OF DRY SAMPLE

(Dryed at 100 Deg. C.)

Phosphoric Acid	0.74	per cent.
Equi. Bone Phosphate of Lime	1.61	" "
Oxide of Iron and Alumina	0.38	" "
Insoluble Siliceous Matter	5.43	" "
Moisture, determined on ~~coarse~~ sample (at 100 Deg. C.)	0.60	" "
Bone Phosphate of Lime, calculated on moist sample	1.62	" "

E. J. Cooper Chemist

L. B. LOCKHART

CONSULTING AND ANALYTICAL CHEMIST

33 1-2 AUBURN AVENUE

ATLANTA, GA.

CERTIFICATE OF ANALYSIS

No. 1/3/1918.

Samples of Ribbon Sugar Cane & Green Sugar Cane.

From Chevelier Corporation, Miami, Fla.

Marked ------	Ribbon Sugar Cane. (purple)	Green Sugar Cane.
Analysis		
Moisture	73.30%	75.20%
Ash	0.43	0035
Protein	0.17	0.20
Crude Fiber	7.55	5.95
Nitrogen-free extract	18.55	18.30
TOTAL	100.00%	100.00%
Total Carbohydrates(Crude Fiber & N-free extract)	26.10%	24.25%
Remarks: Cane Sugar(Sucrose)	10.50%	6.50%
Reducing Sugars	0.35%	1.13%

The above analyses show the sugar content and the feeding value of the samples. It will be noted that the Green Sugar Cane has a lower crude fiber content than the other sample with a correspondingly higher digestibility when used as a feed.

L. B. LOCKHART CONSULTING CHEMIST ANALYTICAL ATLANTA GA.

Respectfully submitted,

L B Lockhart

Chemist.

L. B. LOCKHART
CONSULTING AND ANALYTICAL CHEMIST
33 1-2 AUBURN AVENUE
ATLANTA, GA.

CERTIFICATE OF ANALYSIS

No. 6/5/1919.

Sample Rock #1.

From Chevelier Corporation, Miami, Fla.

Marked (Your letter May 27th).

Analysis

	Per cent.
Moisture	0.25
Silica (SiO_2)	0.06
Phosphates	None
Carbon Dioxide (CO_2)	42.60
Lime (CaO)	54.20
Magnesia (MgO)	0.20
Oxides of Iron & Alumina.	0.36
Alkalies and Undetermined	2.33 2.33
Total	100.00

Remarks:

This analysis shows a limstone of unusual purity. The Silica (sand, etc.), the magnesia, and the iron and alumina being very low. There is present a total of 96.60% Carbonate of Lime. This Limestone would be suitable, chemically, for the manufacture of cement, lime, silica-lime brick, and for agricultural purposes(as ground limestone). Its physical character might be suitable for ballasting and possibly for road building.

Respectfully submitted,

L B Lockhart

Chemist.

CO PY

L O C K H A R T L A B O R A T O R I E S

Atlanta, Ga.

Chevelier Corporation,
Miami, Fla.

Gentlemen:

In connection with the analysis of a pure limestone submitted by us on the 5th of June, I wish to state that this limestone belongs to a special class generally known as "oolite" or oolithic limestone. It is related in its composition and in its method of origin to the form of limestone known as "Travertine".

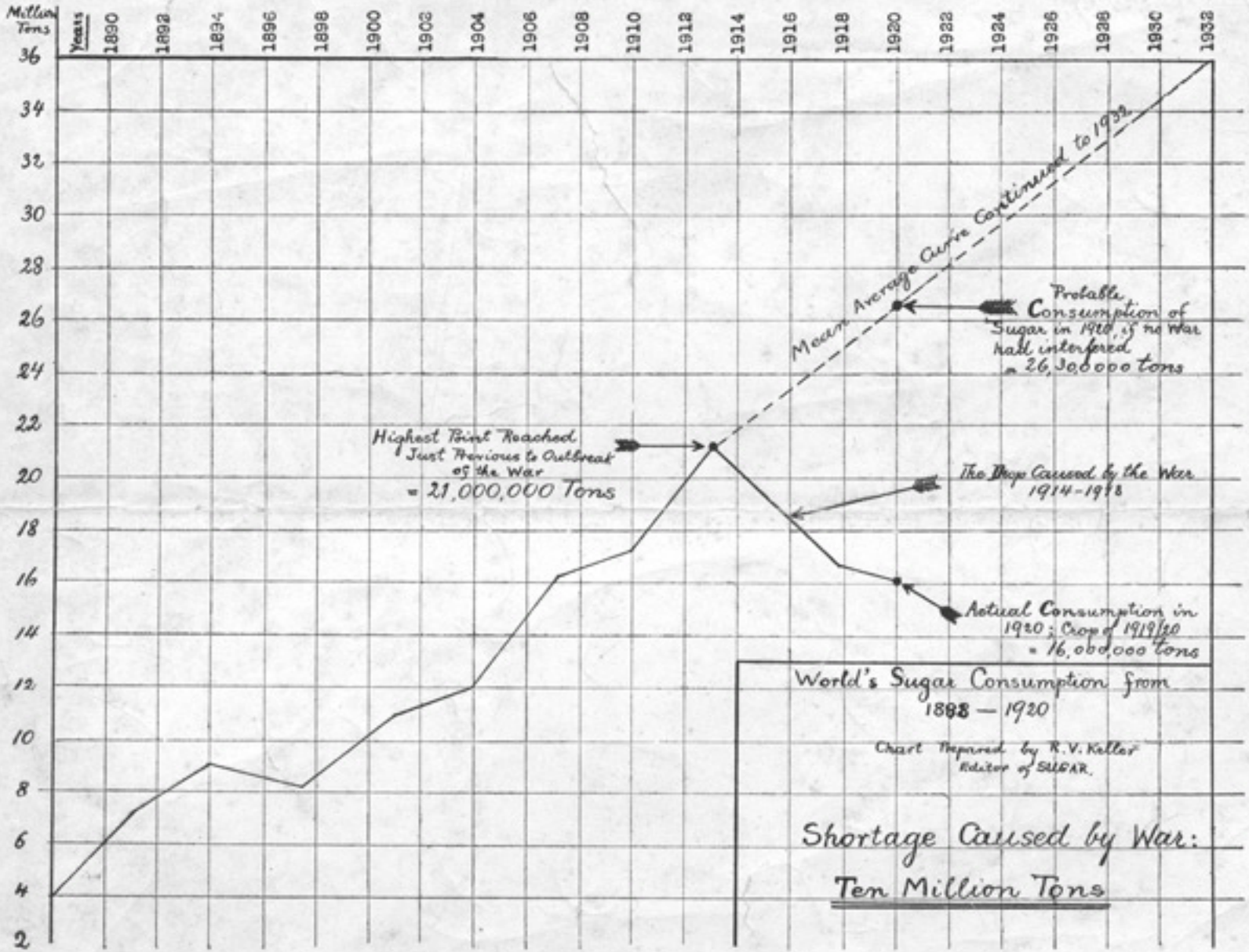
This limestone is probably of comparatively recent origin (geologically speaking). The value of a material of this nature depends largely on the tonnage available, the uniformity, the ease of mining and the location of the deposit with respect to the possible market.

It might have a possible value in certain chemical processes requiring pure lime, such as possibly in purifying sugar.

Yours very truly,

(Signed) LOCKHART LABORATORIES

L/b Per L.B. Lockhart.


Million Tons
Years
36
34
32
30
28
26
24
22
20
18
16
14
12
10
8
6
4
2
1890
1892
1894
1896
1898
1900
1902
1904
1906
1908
1910
1912
1914
1916
1918
1920
1922
1924
1926
1928
1930
1932
Mean Average Curve Continued to 1932
Probable Consumption of Sugar in 1920, if no War had interfered = 26,500,000 tons
Highest Point Reached Just Previous to Outbreak of the War = 21,000,000 Tons
The Drop Caused by the War 1914-1918
Actual Consumption in 1920; Crop of 1919/20 = 16,000,000 tons
World's Sugar Consumption from 1888 — 1920
Chart Inspired by R. V. Keller Editor of SUGAR.
Shortage Caused by War:
Ten Million Tons